

**SCIENCE (BIOLOGY + CHEMISTRY)
ENTRANCE TEST SAMPLE PAPER**

sample paper only provide
10 MCQ and 2 SAQ

Actual Paper
Total 30 MCQ + 4 SAQ

Each MCQ is 2 marks
Each SAQ is 10 marks

Instructions

1. This is a **closed-book** test.
2. It has a time limit of **90 minutes** and allows for only **ONE attempt (submission)**.
3. Alert the invigilator if you are facing technical difficulties.
4. You are to **ensure** that:
 - your laptops, computers and any other devices used for this test is in good functioning order and have uninterrupted power supply and internet connection throughout the duration of the test.
 - you are in a conducive environment throughout the duration of the test.
 - your answers are correctly saved by the end of the test.
5. You are **allowed** to use:
 - a scientific calculator.
 - a blank piece of paper (no larger than A4 size) for rough work. The paper will not be accepted for submission at the end of the test.
6. You are **not allowed** to:
 - leave the test or leave your devices throughout the duration of the test.
 - use the washroom throughout the duration of the test.
 - communicate with any person, either face-to-face or through any communication device, other than the invigilator.
 - refer to any references, e.g. textbooks, resources from a laptop or smart devices etc.
 - share materials (e.g. electronic calculator) during the test.
 - use any communication devices such as mobile phones, tablets, smart watches, headsets during the test.
7. Enter the password provided by the invigilator to start Test paper.

Section A

Choose the most appropriate answer from the options provided. Each MCQ is worth 2 marks.

Biology

Question 1

Which of the following options describes “Diffusion”?

Key: “✓” = True; “x” = False

	Occurs in any substances, e.g., gas and liquid	Takes place through a partially permeable membrane	Substances move down a concentration gradient
A	x	x	✓
B	✓	x	✓
C	✓	✓	x
D	x	✓	✓

Question 2

Figure 1 shows the effect of varying light intensity and CO₂ level on the rate of photosynthesis.

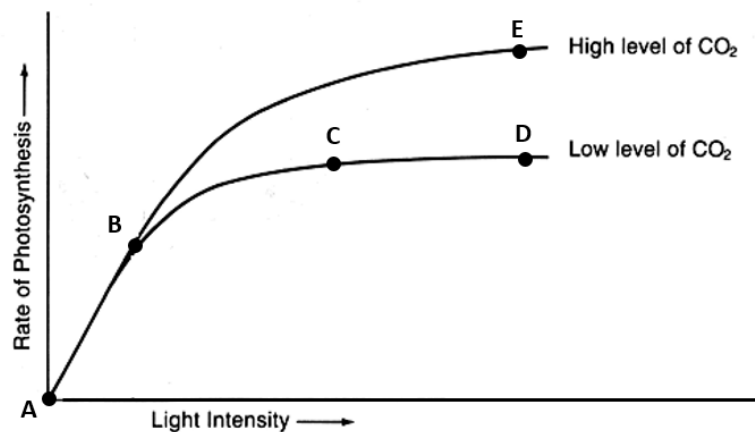


Figure 1

What is limiting the rate of photosynthesis?

- A. Light intensity between Point A to B
- B. Light Intensity between Point C to D
- C. CO₂ level between Point A to B
- D. CO₂ level between Point B to C

Question 3

Which of the following events would directly increase the area of carbon sinks in an ecosystem?

- A. Burning more plants
- B. Humans eating more meat
- C. Increasing soil stability
- D. Draining lakes

Question 4

Which of the following options shows the characteristics of deoxyribonucleic acid (DNA)?

Key: "✓" = True ; "x" = False

	The sugar unit is deoxyribose	It is a double stranded molecule	It is a temporary molecule and is made only when needed	Uracil is one of the nitrogen-containing bases
A	✓	✓	x	x
B	✓	x	✓	x
C	x	✓	✓	x
D	x	✓	x	✓

Question 5

The _____ and _____ of a flowering plant contain haploid nuclei.

- A. pollen and ovum
- B. Ovule and Sepal
- C. Anther and Sepal
- D. Ovum and Ovule

Chemistry**Question 6**

Methanol boils at 65°C and water boils at 100°C. Given that methanol and water are completely miscible with each other, which is the **MOST SUITABLE** method to separate a mixture of these two liquids?

- A. Evaporation
- B. Crystallisation
- C. Fractional distillation
- D. Paper chromatography

Question 7

Two isotopes of carbon are ^{12}C and ^{13}C . Which statement about the isotopes is **TRUE**?

- A. They have the same number of electrons and neutrons.
- B. They have the same number of electrons and protons.
- C. They have the same number of neutrons and protons.
- D. They have the same number of nucleons and electrons.

Question 8

The electronic configuration of atom **D** is 2, 7. The electronic configuration of atom **E** is 2, 6. What is the formula of the compound formed between atoms **D** and **E**?

- A. D_2E
- B. DE_2
- C. D_6E
- D. DE_7

Question 9

A label is missing from a bottle of green solution **C**. In order to identify the solution, two chemical tests are carried out.

Test 1: A few drops of aqueous sodium hydroxide are added to a sample of solution **C**. A green precipitate is formed.

Test 2: Excess aqueous sodium hydroxide and aluminium are added to another sample of solution **C** and heated. A pungent gas, which turns damp red litmus paper blue, is produced.

What is **C**?

- A. Iron(II) nitrate
- B. Iron(III) nitrate
- C. Iron(II) sulfate
- D. Iron(III) sulfate

Question 10

A solution of nitric acid has a concentration of 0.100 mol/dm^3 while a solution of potassium hydroxide has a concentration of 0.125 mol/dm^3 . What is the volume (in cm^3) of potassium hydroxide required to completely neutralize 20.0 cm^3 of nitric acid?

- A. 8.00
- B. 12.0
- C. 16.0
- D. 32.0

End of Section A

Section B

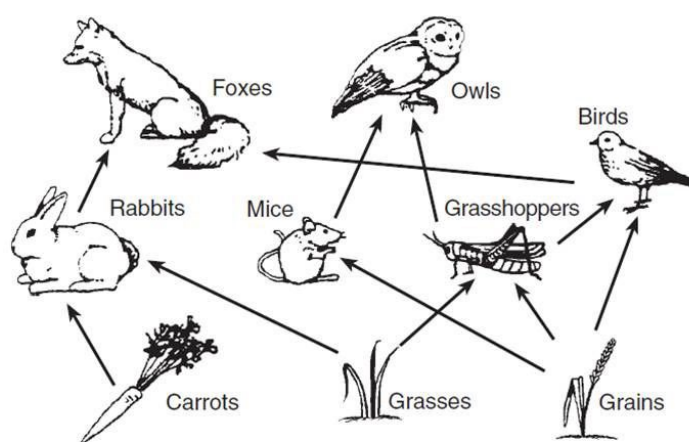
Provide your answers in the spaces below each question.

NOTE: Round off your answers to 2 decimal places, if applicable.

Biology (Total marks: 10 marks)

Question 11

Figure 2 shows a food web. Answer the following questions.



(Not drawn to scale)

Figure 2

a) Identify **ONE** producer and explain why it is a producer in the food web. (4 marks)

b) Identify **TWO** secondary consumers and explain why they are secondary consumers in the food web. (3 marks)

c) State **THREE** ways in which energy may be lost between trophic levels. (3 marks)

Chemistry (Total marks: 10 marks)

Question 12

An atom of an element **L** has one electron in its second electron shell.

- a) State the atomic number of this element. (1 mark)

- b) State which group and period of the periodic table this element is in. (2 marks)

- c) What is the name of this element? (1 mark)

- d) Identify **TWO** other elements in the same group. (2 marks)

- e) Explain why this element has similar chemical properties as other members of its group in the periodic table. (1 mark)

- f) Element **L**, oxygen and fluorine are in the same period.
 - (i) Explain why these three elements are in the same period. (1 mark)
 - (ii) Write the name of the compounds formed between: (2 marks)
 - Element **L** and oxygen:
 - Element **L** and fluorine:

End of Section B

Periodic Table

The Periodic Table of the Elements

		Group												
I	II	III	IV	V	VI	VII						0		
7 Li lithium 3	9 Be beryllium 4	1 H hydrogen 1										4 He helium 2		
23 Na sodium 11	24 Mg magnesium 12	11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10	27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18	
39 K potassium 19	40 Ca calcium 20	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Ni nickel 28	64 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	84 Kr krypton 36
85 Rb rubidium 37	88 Sr strontium 38	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	101 Ru ruthenium 44	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	127 I iodine 53	131 Xe xenon 54	
133 Cs caesium 55	137 Ba barium 56	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	210 Po polonium 84	222 Rn radon 86
87 Fr francium	88 Ra radium											89 Ac actinium		

*58-71 Lanthanoid series											
140 Ce cerium 58	141 Pr praseodymium 59	144 Nd neodymium 60	150 Sm samarium 62	152 Eu europium 63	157 Gd gadolinium 64	162 Dy dysprosium 66	165 Ho holmium 67	167 Er erbium 68	169 Tm thulium 69	173 Yb ytterbium 70	175 Lu lutetium 71

†90-103 Actinoid series													
232 Th thorium 90	238 U uranium 92	238 Pa protactinium 91	238 Np neptunium 93	238 Pu plutonium 94	238 Am americium 95	238 Cm curium 96	238 Bk berkelium 97	238 Cf californium 98	238 Es einsteinium 99	238 Fm fermium 100	238 Md mendelevium 101	238 No nobelium 102	238 Lr lawrencium 103

Key

a	X
b	

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

END OF PAPER